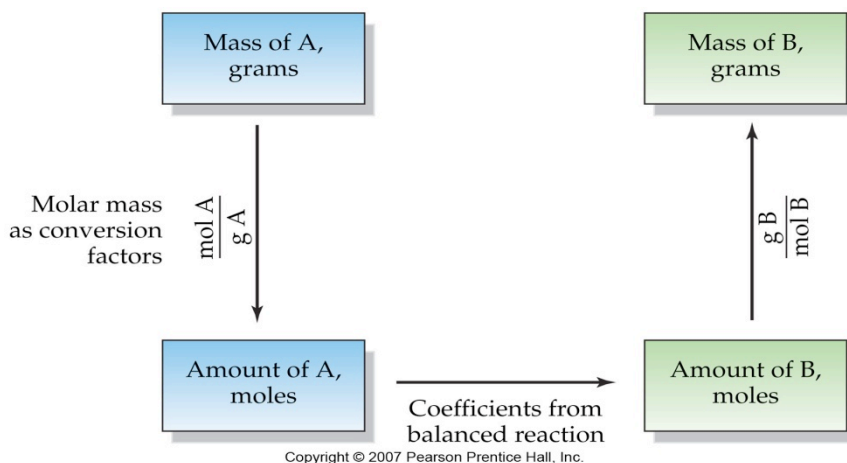
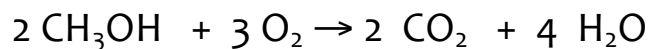


How to use reaction stoichiometry

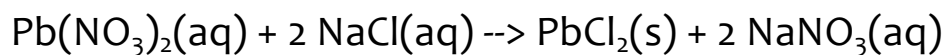
1. Write and balance the chemical equation for the reaction.
2. Determine molar masses of substances involved in the calculation.
3. Use the coefficients of the balanced equation to convert the moles of the given substance to the moles of the desired substance.
4. Use the molar mass to convert the moles of the desired substance to grams of the desired substance.

Examples:

1. If we react 100 grams of CH_3OH , how many grams of H_2O are formed?



2. Consider the following reaction:



If you have 5 moles of $\text{Pb}(\text{NO}_3)_2$, how many moles of PbCl_2 can you make?

If you have 5 moles of NaCl , how many moles of PbCl_2 can you make?