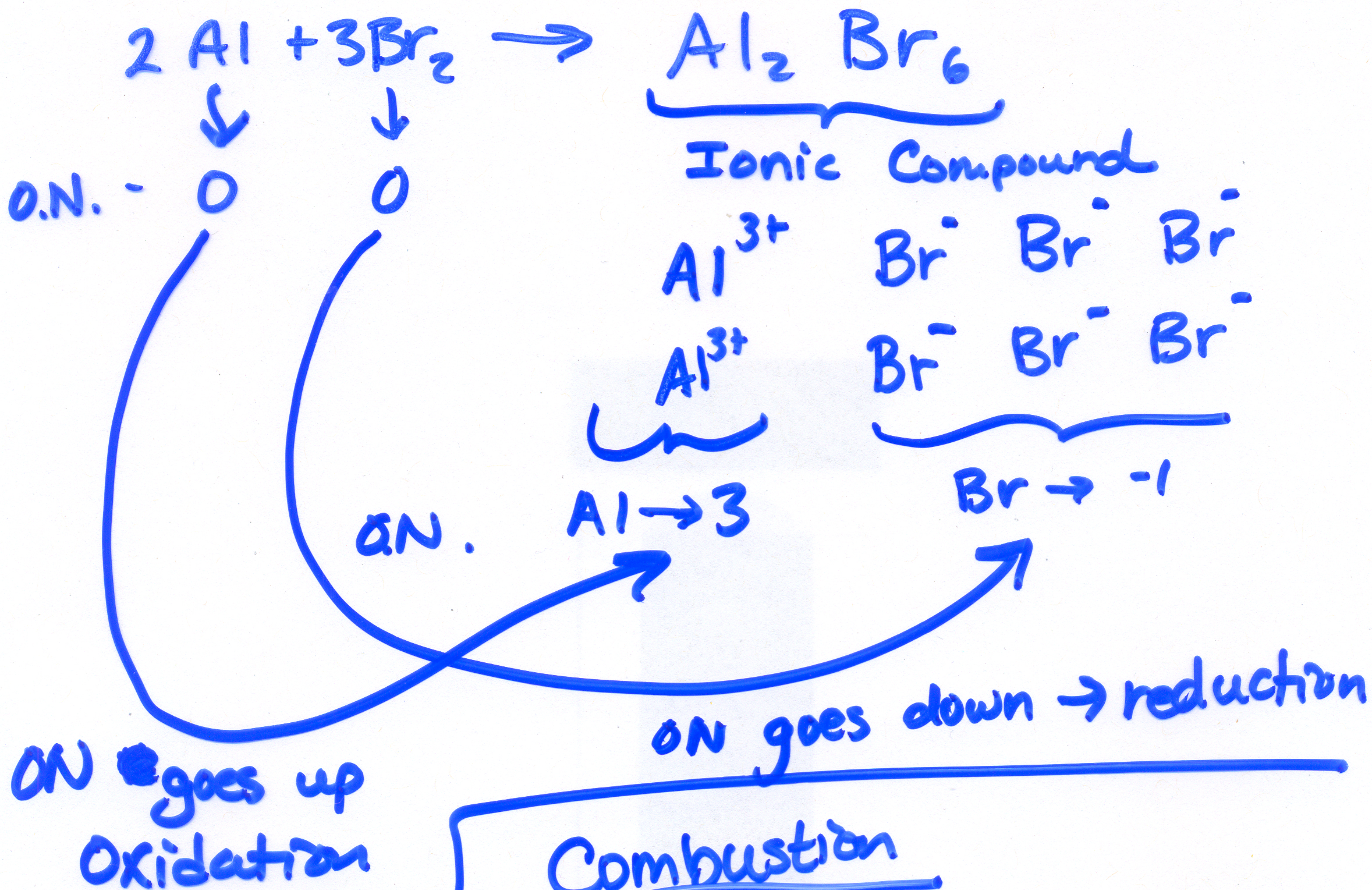
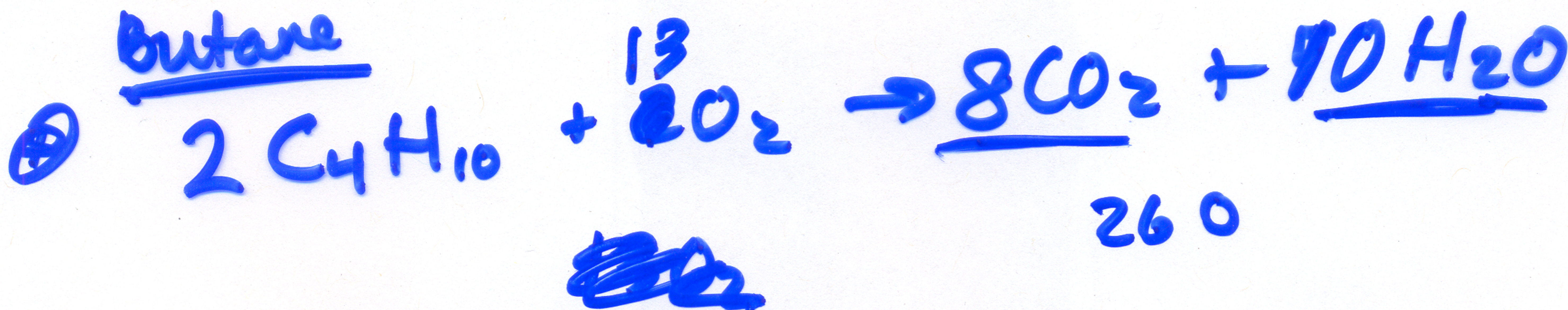
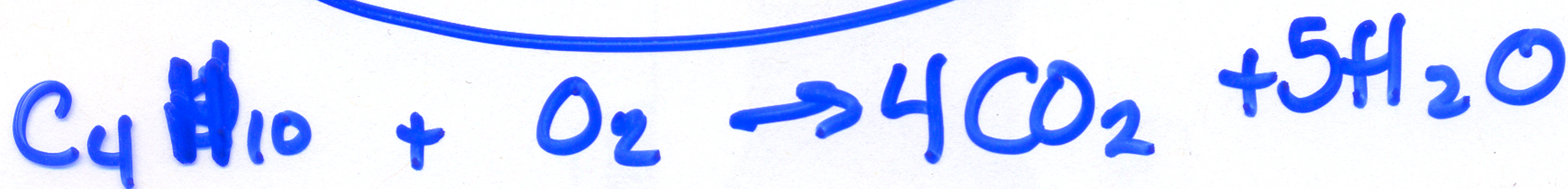
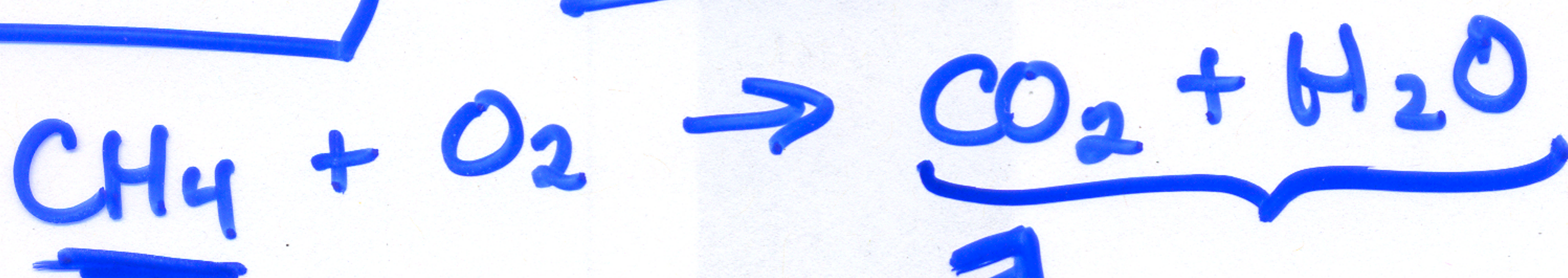
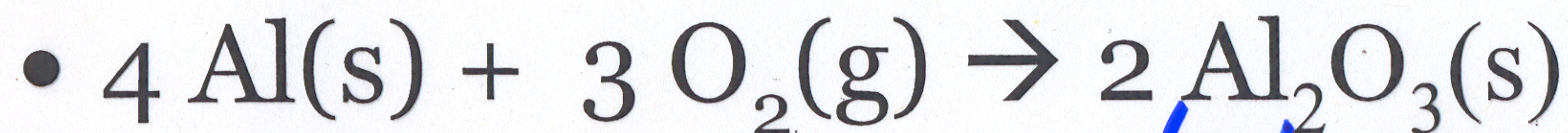
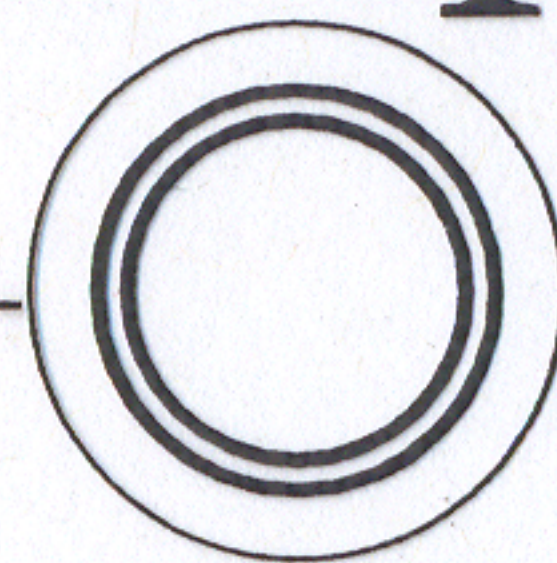




Combustion



Examples



ON:

0

0

Net +6

(-6)

Al
+3

Al
+3

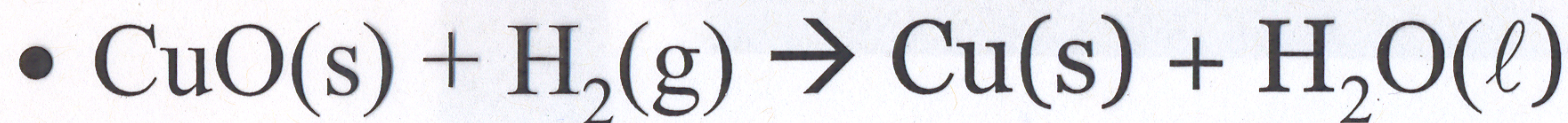
increase
= oxidation

Add O $\Rightarrow$ Redox

O is reduced

Al is oxidized

@ O has ON of -2



ON: +2

-2

0

0

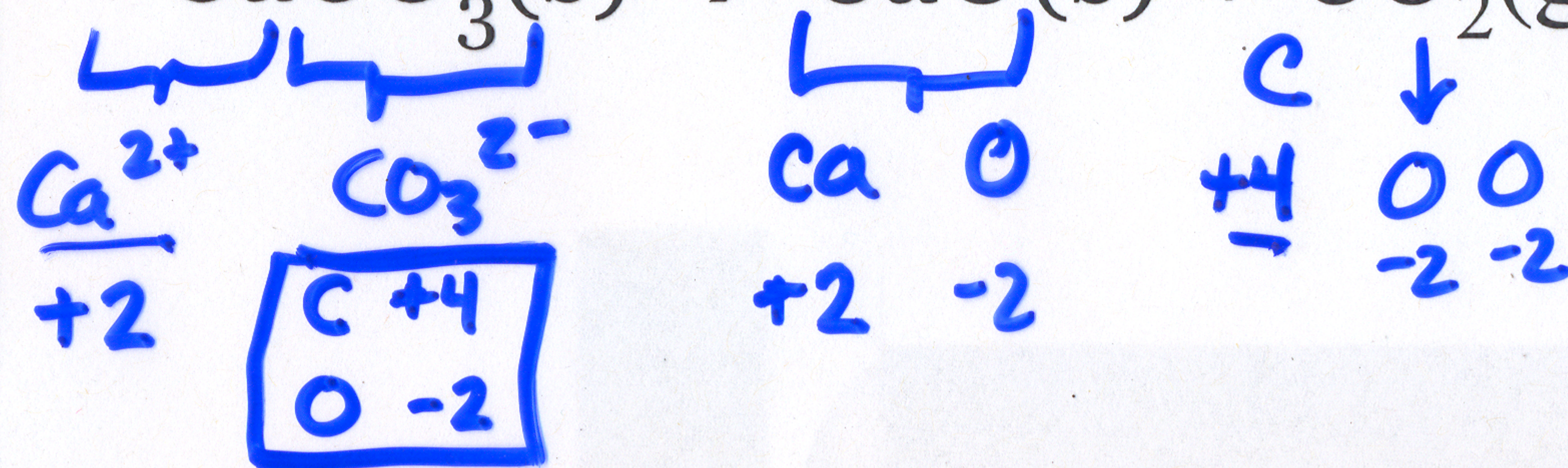
+1

-2

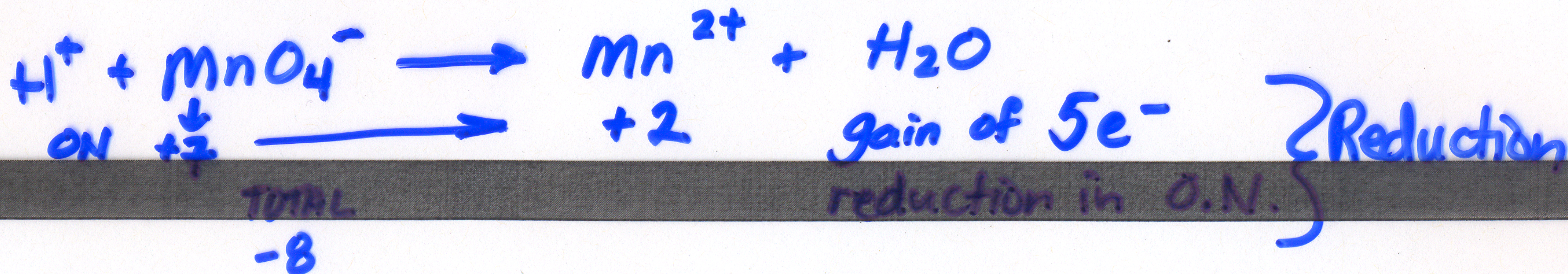
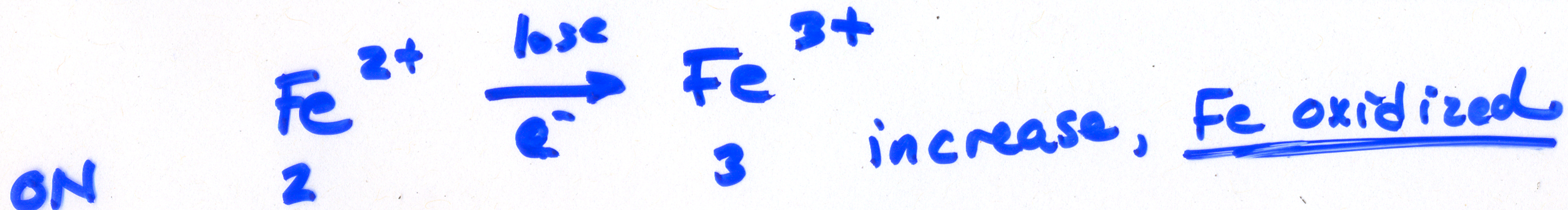
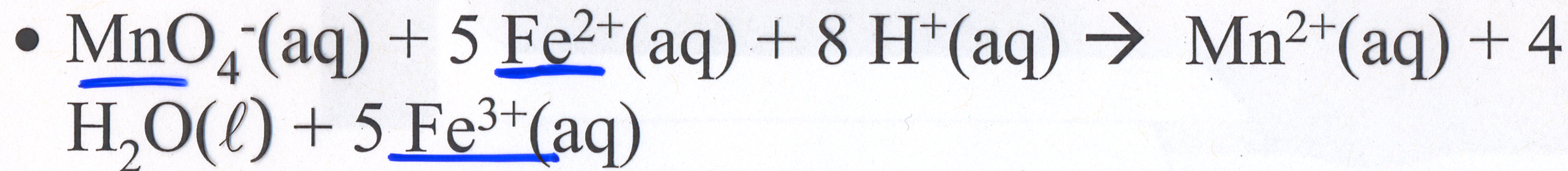
+1

Cu $\rightarrow$ O.N. decreased, loss of O $\rightarrow$ reduced (Oxidizing agent)
 H $\rightarrow$ O.N. increased, gain of O $\rightarrow$ oxidized (Reducing agent)
 O $\rightarrow$ O.N. stays same

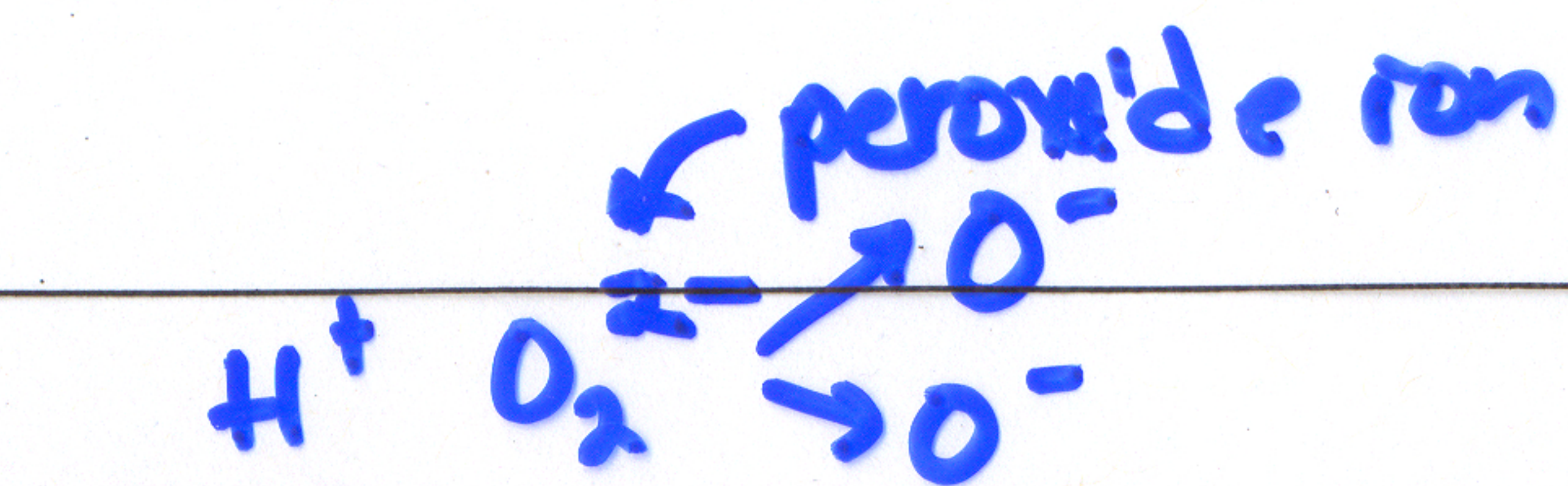
Examples



Not redox
 (NO O.N.'s changed)



Hydrogen Peroxide

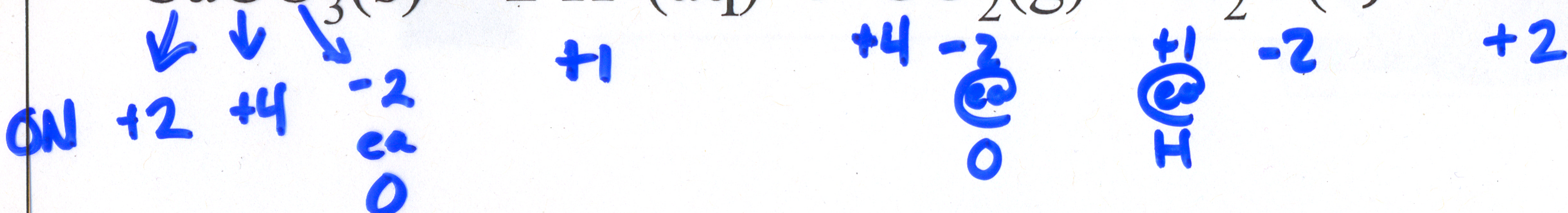
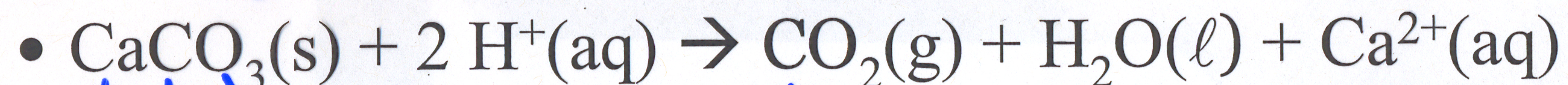
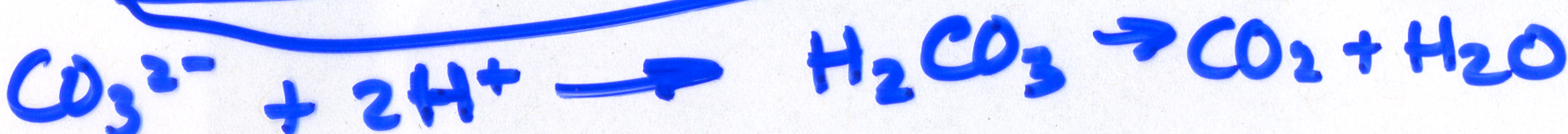


Examples

Decomposition

O is reduced to H₂O
O is oxidized to O₂

Redox



No O.N.'s changed → Not Redox

Acid-Base → Gas Forming