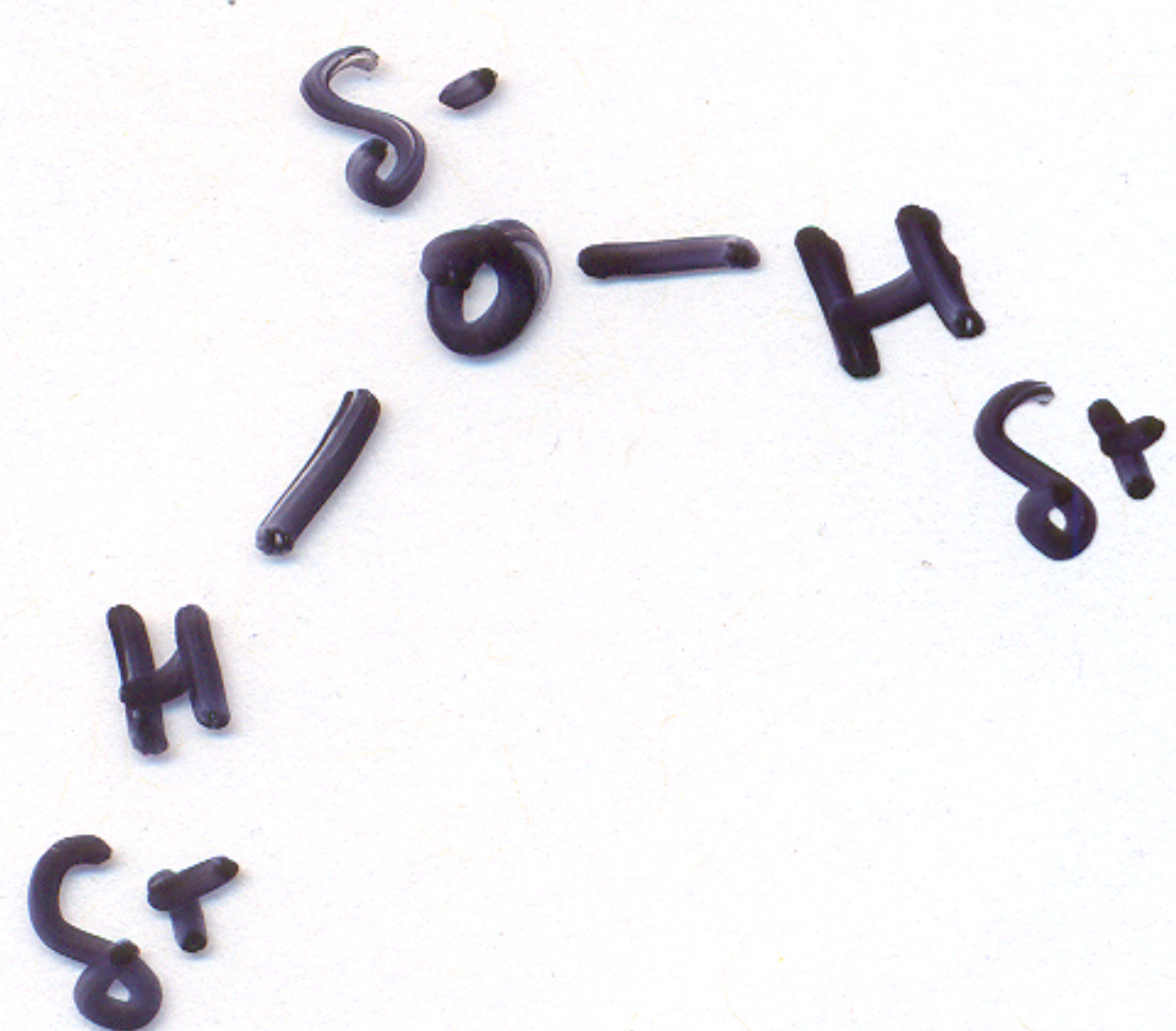
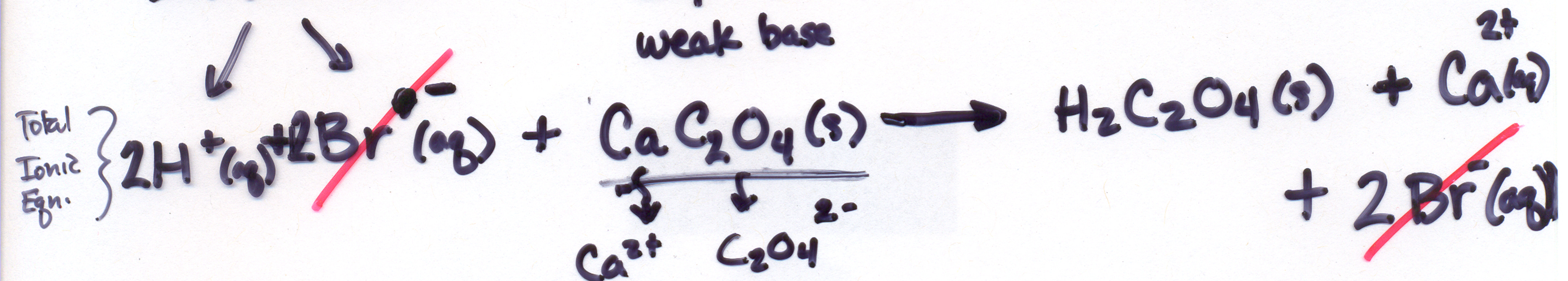
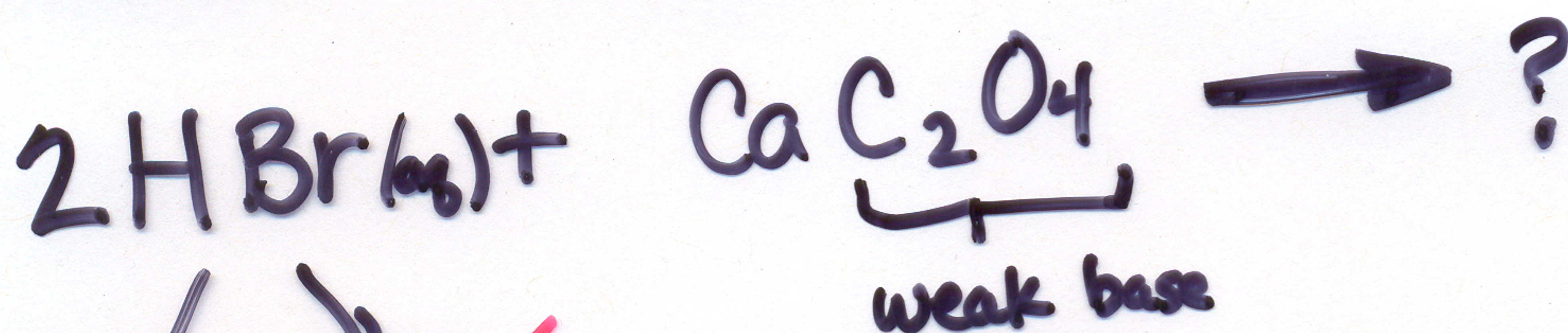
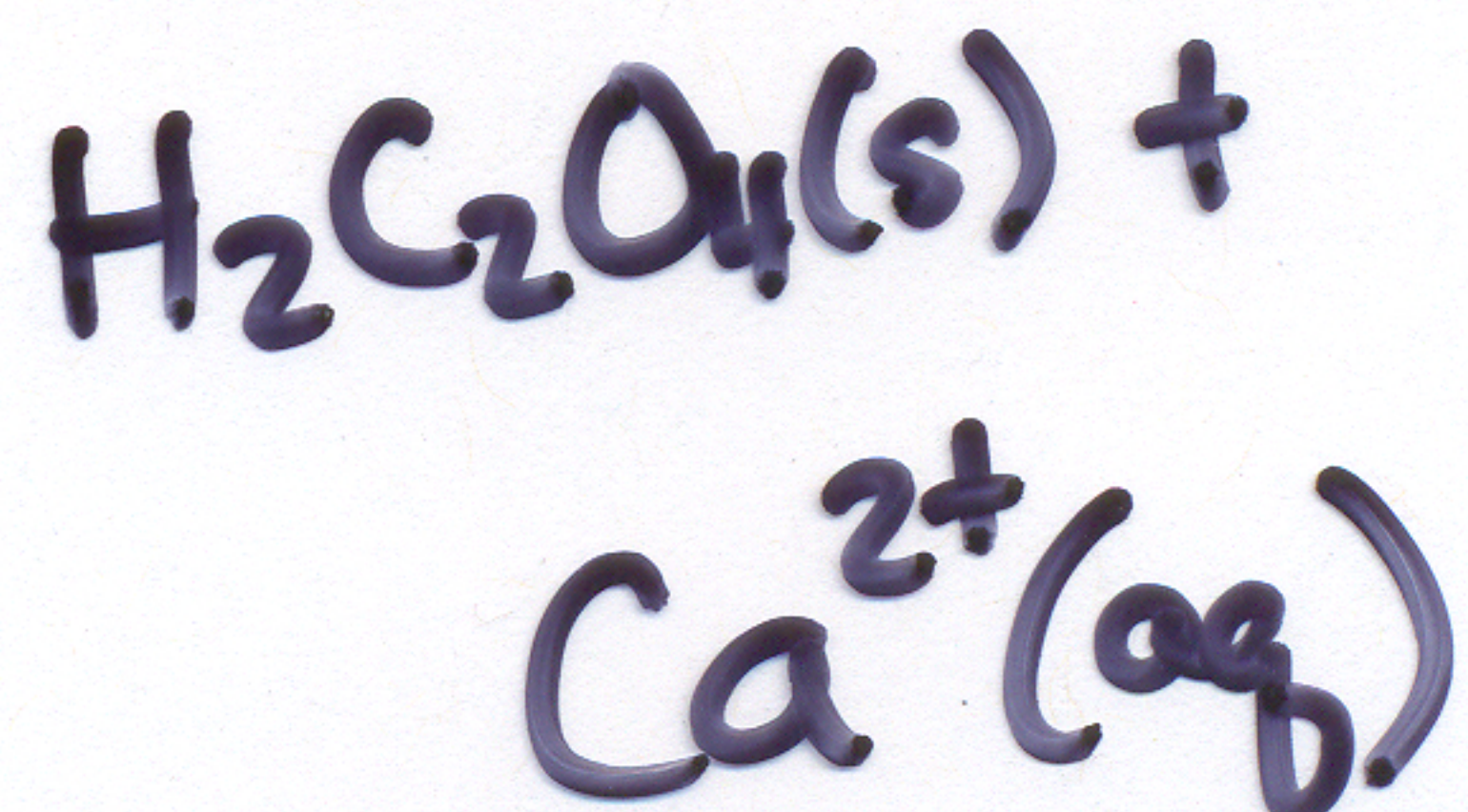


Write a net ionic equation for the reaction:

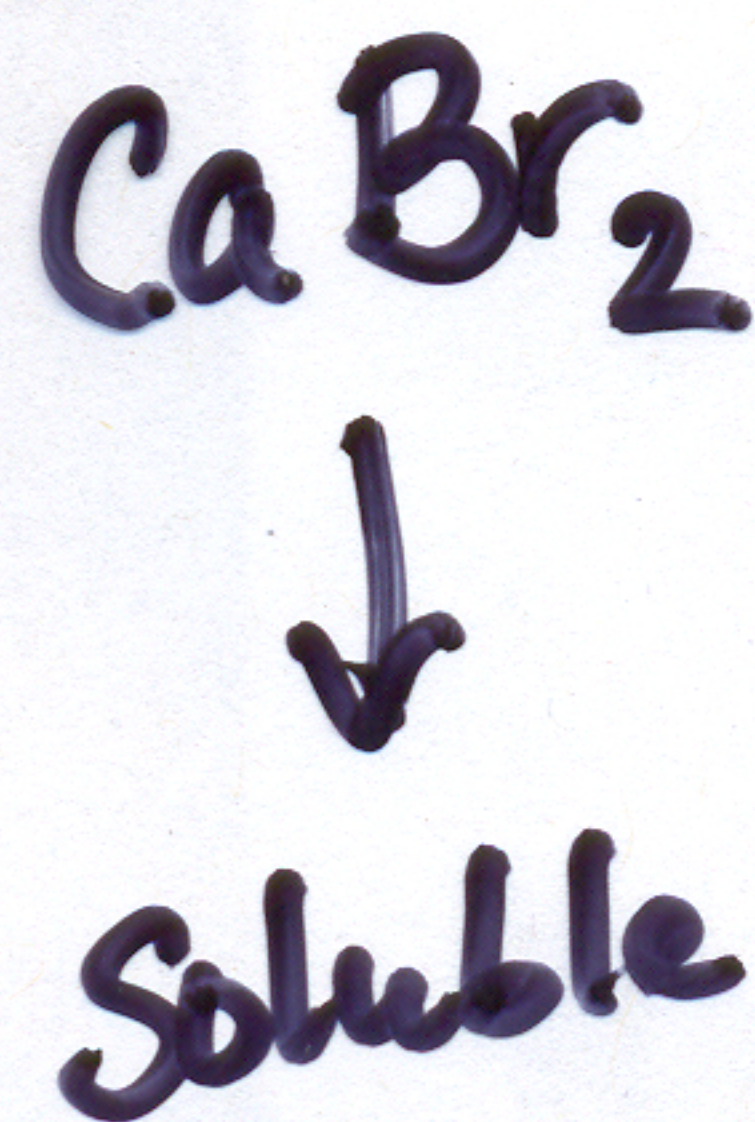
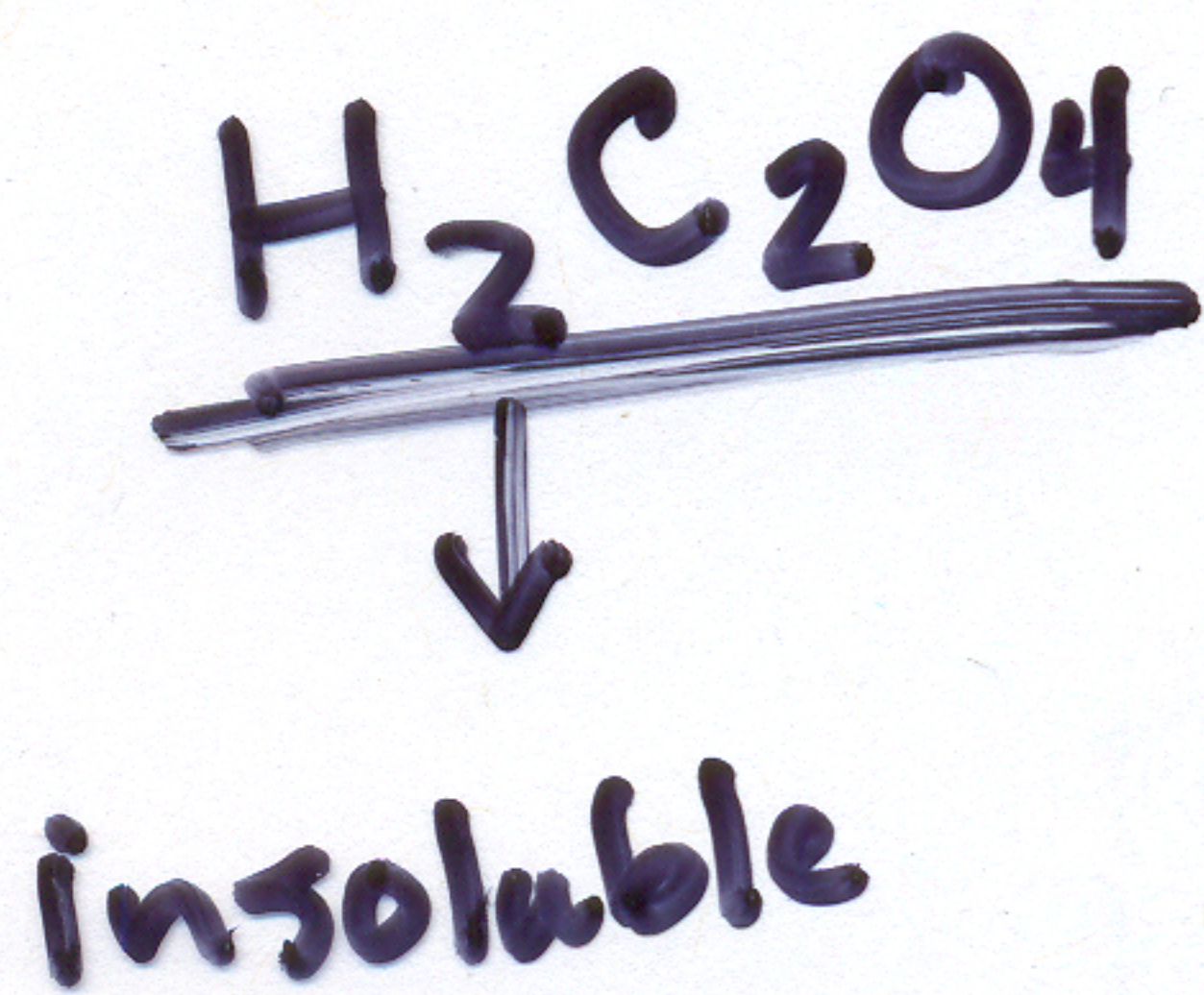
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Net ionic Equation: $2\text{H}^+(aq) + \text{CaC}_2\text{O}_4(s) \rightarrow$



Two possible products:

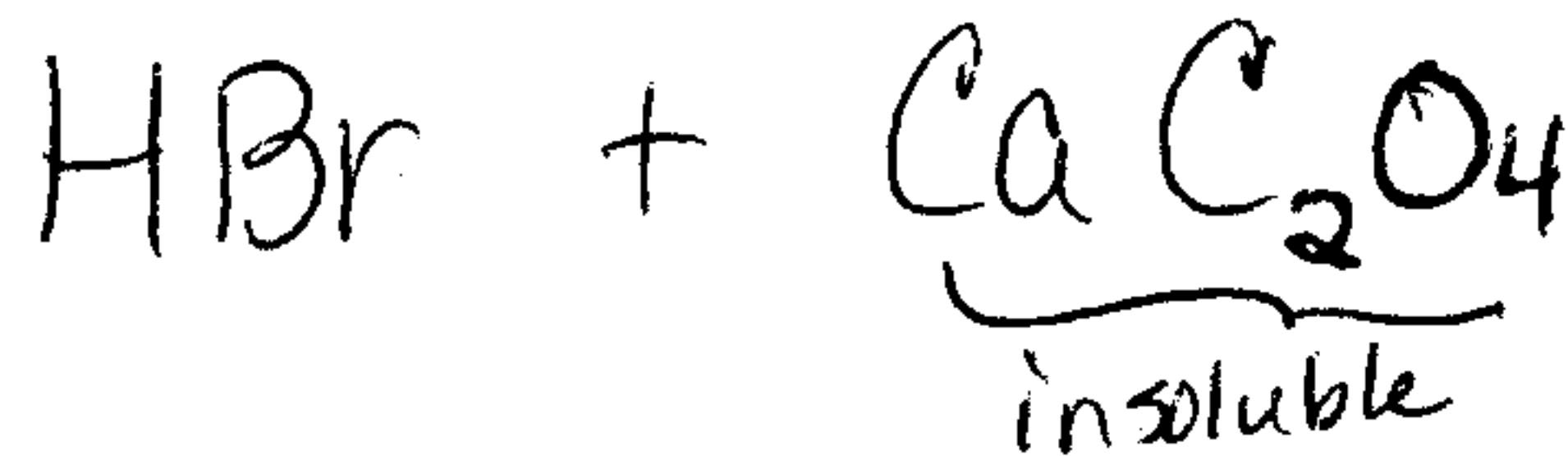


Bromines are spectator ions.

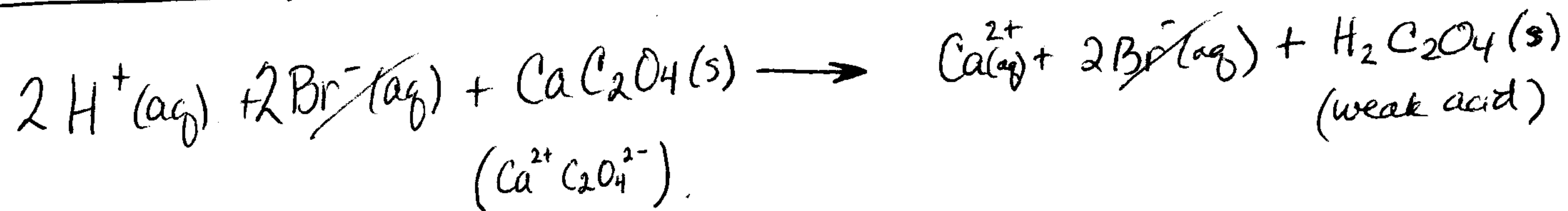
H_3PO_4 weak acid
in water $\rightarrow$ stay as H_3PO_4

Strong Acid

$\text{C}_2\text{O}_4^{2-}$ = weak base (conjugate)



Total Ionic Equation:



Net Ionic Equation:

