Clicker Questions

What is the pH when $[H_3O^+] = 0.045$ M?

What is the pH when $[OH^-] = 6.5 \times 10^{-4}$ M?

What is the $[H_3O^+]$ when pH = 3.66 M?

What is the $[OH^-]$ when pH = 5.84 M?
\[ [H_3O^+] = 10^{-pH} \]

\[ pH = -\log[H_3O^+] \]

\[ pOH = -\log[OH^-] \]

\[ [H_3O^+][OH^-] = 10^{-14} \]

\[ pH + pOH = 14 \]