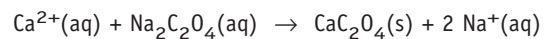


PRE-LABORATORY ASSIGNMENT
PERCENT COMPOSITION OF A MIXTURE

Name _____

Section _____

The amount of calcium ion present in milk can be determined by adding oxalate ion, $\text{C}_2\text{O}_4^{2-}$ (in the form of its sodium salt $\text{Na}_2\text{C}_2\text{O}_4$). The insoluble compound calcium oxalate, CaC_2O_4 , is precipitated.



Suppose you have a 75.0-g sample of milk and isolate 0.288 g of CaC_2O_4 from it. What is the mass percentage of calcium in the milk?