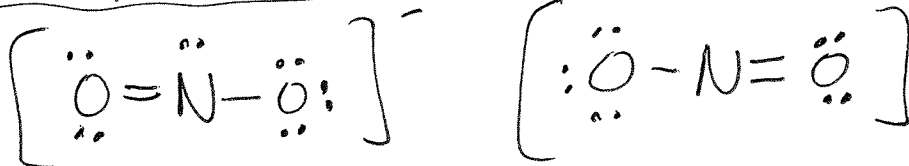
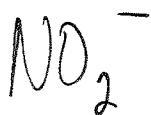
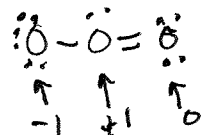




Central O uses sp² hybrids

epg = trigonal planar
 mg = bent
 bond angle = ~120°
 non-polar

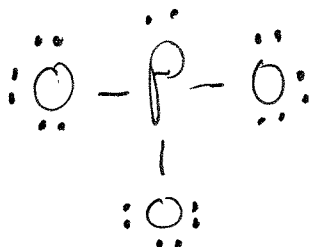
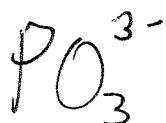
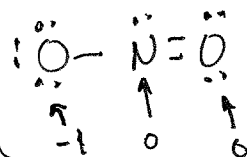
formal charges



N uses sp² hybrids

epg = trigonal planar
 mg = bent
 bond angle = ~120°
 polar

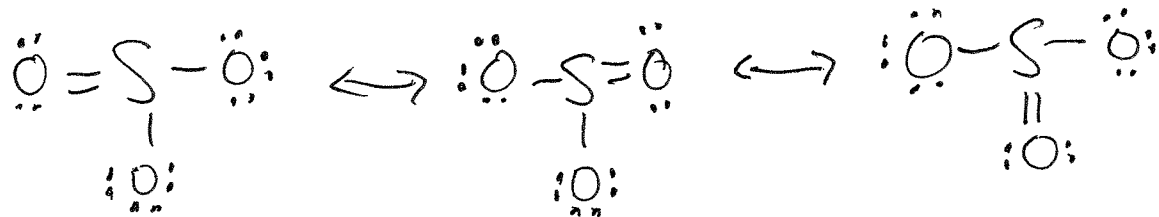
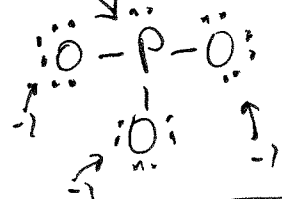
formal charges



P uses sp³ hybrids

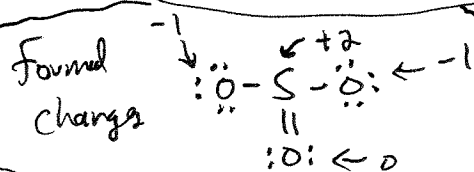
epg = tetrahedral
 mg = trigonal pyramidal
 bond angle = ~109.5°
 polar

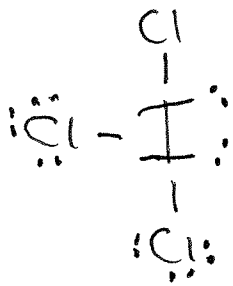
formal charges



epg = trigonal planar
 mg = trigonal planar
 bond angle = 120°
 non-polar

S uses sp² hybrids





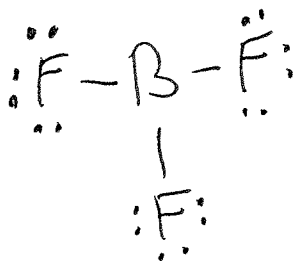
epg = trigonal bipyramidal

mg = T-shaped

bond angles = 90° , 180°

polar

I uses sp^3d
hybrid orbitals



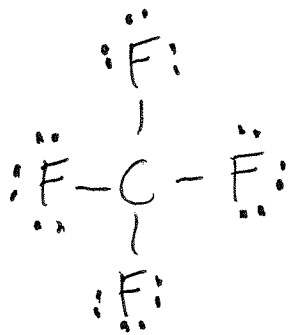
epg = trigonal planar

mg = trigonal planar

bond angles = 120°

non-polar

B uses sp^2
hybrids



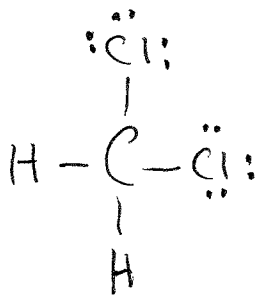
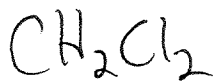
epg = ~~trigonal plan~~ tetrahedral

mg = tetrahedral

bond angles = 109.5°

non-polar

C uses
 sp^3 hybrids



epg = tetrahedral

mg = tetrahedral

bond angles = 109.5°

polar

C uses
 sp^3 hybrids