

Name: _____

Section: _____

Prelab Assignment: Determination of an Unknown Acid

- To standardize a solution is to:
 - Heat it.
 - Add a standard to it.
 - Determine its exact concentration.
- According to the balanced chemical equation one mole of NaOH reacts with one mole of oxalic acid.

True False

- Using the information in part b) of the experimental procedure, determine the approximate concentration of the NaOH solution you will prepare.

**Hint: you can use the quick formula $C_1V_1 = C_2V_2$
(Concentration of solution 1 times volume of solution 1 = concentration of solution 2 x volume of solution 2; solve for C_2)**

- In order to determine the mass of oxalic acid dihydrate needed to standardize the NaOH solution you will prepare you will need to know:
 - The **approximate** concentration of the NaOH
 - The mole ratio for the reaction between oxalic acid and NaOH
 - The molar mass of oxalic acid
 - Only A and C are required
 - A, B and C are required
- Part c) of the procedure states that you will have calculated the mass of $H_2C_2O_4 \cdot 2H_2O$ needed to standardize the NaOH. Using information from other questions in this quiz, and assuming you would like to use about **25 mL of your NaOH solution**, calculate the mass of oxalic acid dihydrate needed to standardize the NaOH.

- What is your MeasureNet ID number?